
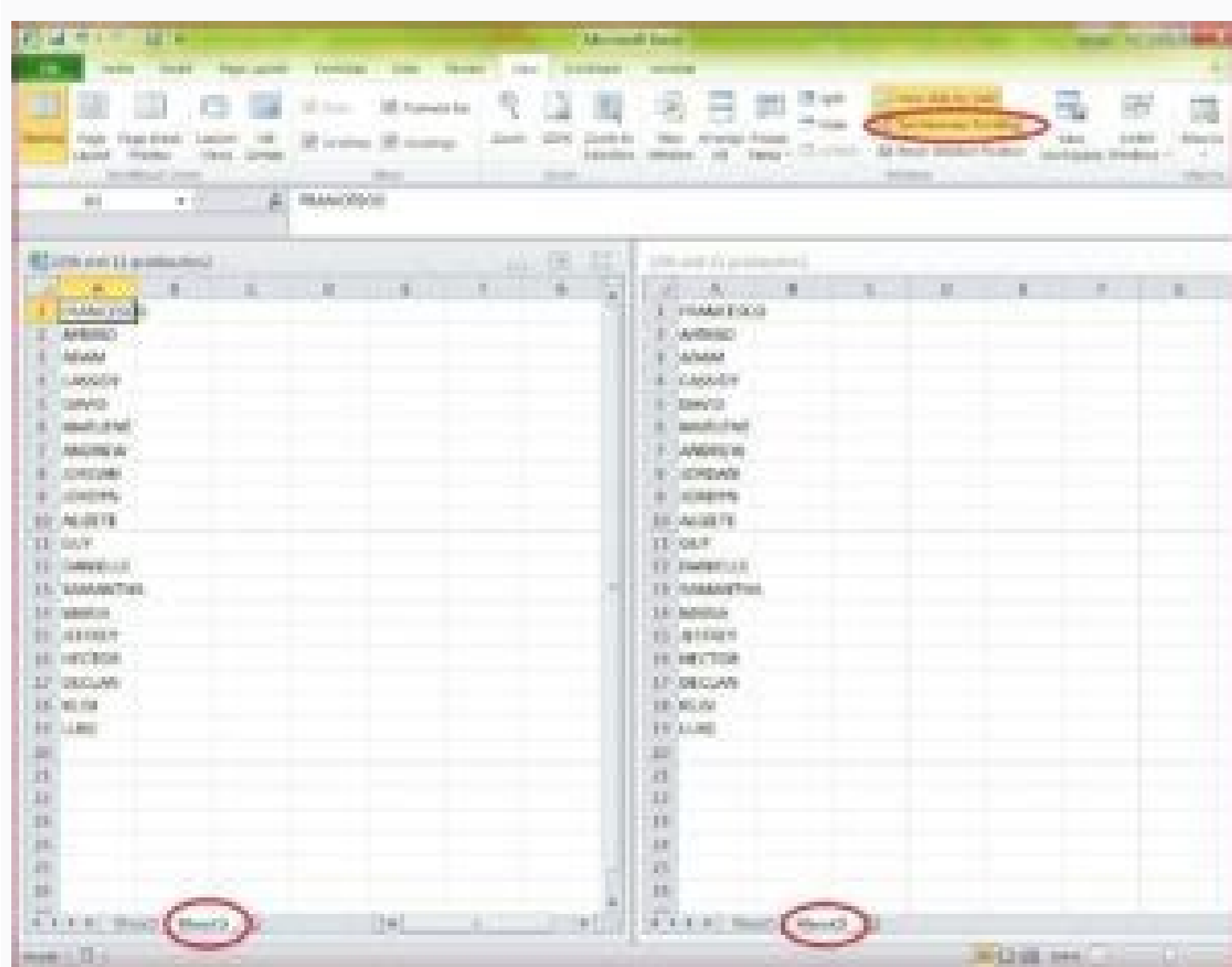


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4.3 ~ Balancing Chemical Equations

skeleton equation = formulas of the reactants & products only, no balancing of # of atoms

- due to the Law of Conservation of Mass, mass of reactants = mass of products
 - we cannot change the formulas therefore we must change the number of molecules
 - we do this by adding coefficients in front of the formulas

coefficient = number written in front to the chemical symbol or formula

balanced chemical equation = when reactants & products contain an equal number of atoms of each type

- by balancing an equation, the mass of the reactants will be equal to the mass of the products

How to Balance an Equation

Ex: Solid iron reacts with oxygen gas to form iron (III) oxide

Step 1: Write the word equation for the reaction

Step 2: Write the skeleton equation by replacing each word with the correct formula

Step 3: Count the number of atoms of each type in reactants and products

Step 4: Multiply each of the formulas by the appropriate coefficients to balance the number of atoms

- Hint: Start with compounds 1st, then elements

Practice problems p. 165 #1, 2

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